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Electrochemistry

NEET/JEE main crash course

① Electrochemical cell

Salt Bridge & its funcⁿ # KNO_3, NH_4NO_3 etc

Inert electrolyte

+
paste of polysaccharide
(agar-agar)

* KCl is not used as electrolyte in
Ag⁺/Ag, Pb, Hg

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Electrochemistry

① Electrochemical cell

a for Galv



Given $E^\circ_{\text{cell}} = +0.44 \text{ V}$

$+0.34 \text{ V}$

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ΔG° and K_{eq} for Galvanic cell

① $\Delta G^\circ = -nFE^\circ_{\text{cell}}$ ② $\Delta G^\circ = -RT \ln K_{eq}$

$n = \text{no. of } e^- \text{ involve}$

$F = 96500 \text{ C}$

$R = 8 \text{ Cal} = \frac{25}{3} \text{ J}$

③ find ΔG° & K_{eq} for Daniell cell

$E^\circ_{\text{cell}} = 1.1 \text{ V}$

$\Delta G^\circ = -2$

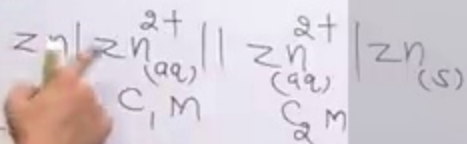
Electrochemistry

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① Electrochemical cell

Concⁿ cell ÷ Same material is present with diff concⁿ or pressure

$$E_{\text{cell}}^{\circ} = 0$$



$$E_{\text{cell}} = \frac{0.06}{n} \log \left(\frac{C_2}{C_1} \right)$$

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② Electrochemical series

<u>Q</u>	Metal	$E^\circ_R (M^{n+}/M)$
	Li	-3.05 V
	Ba	-2.9 V
	Na	-2.71 V
	Mg	-2.37 V

compare reducing power

Li >

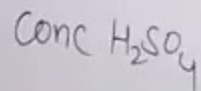
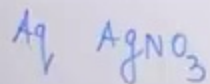
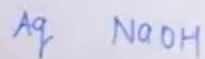
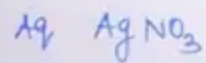
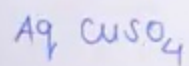
Q → Compare reactivity

Ba, Zn, K, Cu

Electrochemistry

③ Electrolytic cell

Electrolyte

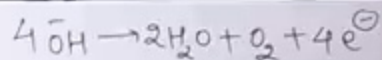


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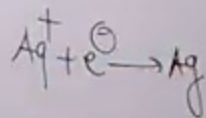
Electrode
Pt or graphite

"

product at cathode



product at anode



Electrochemistry

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- ④ Resistance, conductance, resistivity
Conductivity of a cell.

① Resistance (R)

$$R = \frac{\rho l}{A} \quad \Omega \text{ (ohm)}$$

ρ = Resistivity or sp. resistance

l = Sep. b/w electrodes (cm)

A = Area of crosssection (cm²)

$$\rho = \frac{RA}{l} \quad \text{ohm cm}$$



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